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ESTIMATION OF SODIUM CHLORIDE

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NaCl is assayed by using silver nitrate

- **Take about 100 ml of water in a cleaned and dried 1000 ml volumetric flask.**
- **Add about 17 gm of Silver Nitrate with continues stirring.**
- **Add more about 700 ml of distilled water, mix.**
- **Make up the volume 1000 ml with water. Mix solution thoroughly.**
- **Keep the solution for at least one hour and then carry out the standardization**



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SILVER NITRATE SOLUTION STANDARDIZATION

- **Transfer about 50 mg, accurately weighed, of reagent-grade sodium chloride, previously dried at 110° for 2 hours, to a 150-ml conical flask. Dissolve in 5 ml of distilled water. Add 2.5 ml of acetic acid, 25 ml of methanol, and about 0.25 ml of Eosin Y indicator.**



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- **ANALYSIS:** Dissolve the Sample in 50 mL of water, and titrate with 0.1 N silver nitrate.
- determining the endpoint potentiometrically.
- Each mL of 0.1 N silver nitrate is equivalent to 5.844 mg of sodium chloride (NaCl).
- Acceptance criteria: 99.0%–100.5% on the dried basis



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THANK YOU