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STANDARDISATION OF OXALLIC ACID AND SODIUM HYDROXIDE

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STANDARDISATION OF OXALLIC ACID

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It is done by using NaOH using phenolphthalein as an indicator,

Preparation of NaOH

- **0.1 M solution Weigh 2 g of NaOH pellets in a clean, preweighed 100-mL beaker on a balance.**
- **Put 250 mL of distilled water into a clean 500 mL volumetric flask. Add the NaOH pellets completely.**
- **When the NaOH has dissolved completely add another 250 mL of distilled water up to the marking. Mix thoroughly.**



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Standardization oxalic acid of solution)

Place the standard oxalic acid solution 0.025 M in the burette.

Transfer 10 mL of the NaOH solution into a 125 mL conical flask with the aid of a pipette, add 1-2 drops of phenolphthalein indicator and titrate with the standard oxalic acid. Repeat the titration until duplicate determinations agree within 0.05mL of each other. (If three titrations do not result in the desired precision, it will be necessary to conduct additional titration.



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Record all the titration volume in the table and calculate,

- **i) Molarity of prepared oxalic acid solution,**
- **(ii) Normality of oxalic acid solution**
- **(iii) moles of in oxalic acid 500mL,**
- **(iv) mass of oxalic acid in 500mL**
- **(v) % Purity of oxalic acid**



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STANDARDISATION OF SODIUM HYDROXIDE

This is done by using oxalic acid as primary standard and phenolphthalin as indicator as the same method as of oxalic acid.



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THANK YOU