

See discussions, stats, and author profiles for this publication at: <https://www.researchgate.net/publication/335338544>

How to prepare different types of buffer solutions- An overview

Method · August 2019

CITATIONS

0

READS

86,627

1 author:



[Tridip Boruah](#)

Gauhati University

20 PUBLICATIONS 88 CITATIONS

SEE PROFILE

EXPERIMENT NO: 1

DATE:

PREPARATION OF BUFFER

AIM:

To prepare the buffer at required pH.

PRINCIPLE:

The pH meter measures at electrical potential developed by pair of electrode pins in a solution. For measurement of pH, an electrode system sensitive to change in H⁺ ion concentration of solution is taken. The electrode system consists of sequence of electrode whose potential raise with pH (H⁺ concentration of the solution).

PROCEDURE:

1. ACETIC ACID- SODIUM ACETATE BUFFER:

REAGENTS REQUIRED:

Acetic Acid 0.2M: 1.5 ml of glacial acetic acid is made upto 100ml with distilled water.

Sodium Acetate Solution: 0.64 gm of sodium acetate or 2.72gm of sodium acetate trihydrate is dissolved in 100ml Distilled water.

PROCEDURE:

Pipette out exactly 36.2ml of sodium acetate solution into 100ml of standard flask and add 14.8ml of glacial acetic acid, make the volume 100ml using distilled water using distilled water. This gives 0.2 M of acetic acid and sodium acetate buffer. The pH is measured with pH meter.

The pH meter is first standardized with pH buffer. Wash electrode with distilled water and introduced into 0.2M acetic acid-sodium acetate buffer prepared, the pH of solution is 4.6.

RESULT:

36.2ml Sodium acetate and 14.8 ml glacial acetic acid were mixed and buffer was prepared. pH was measured initial reading observed was 4 which made upto 4.6 with 5N NaOH.

2. BARBITONE BUFFER:

REAGENTS REQUIRED:

- Diethyl barbituric acid.
- Sodium diethyl barbiturate

PROCEDURE:

Dissolve 2.85gm of diethyl barbituric acid and 14.2gm of sodium diethyl barbiturate in distilled water and upto 1 liter. This gives the barbitone buffer.

The pH meter is first standardized with pH buffer. Wash electrode with distilled water and introduced into barbitone buffer prepared, the pH of solution is 6.8.

3. CITRATE BUFFER:

REAGENTS REQUIRED:

- Citric acid: Dissolve 2.101 gm of citric acid in 100ml distilled water.
- Sodium citrate solution 0.1 M: Dissolved 2.941gm of sodium citrate in 100ml distilled water.

PROCEDURE:

46.5ml of citric acid with 3.5ml of sodium citrate solution and upto 100ml with distilled water. It corresponds to 0.1 M citrate buffer and standardised with pH meter and measures the pH of the prepared solution. This gives citrate buffer at pH 2.5.

RESULT:

Citrate buffer was prepared and the pH observed was 4.8 which was adjusted to 2.5 using 1N HCl and 5N NaOH.

4. CARBONATE- BICARBONATE BUFFER:**REAGENTS REQUIRED:**

- Sodium carbonate solution 0.2M: Dissolve 2.12gm of anhydrous sodium carbonate in 100ml Distilled water.
- Sodium bicarbonate solution: Dissolve 1.68gm of sodium bicarbonate in 100ml of distilled water.

PROCEDURE:

Pipette out exactly 27.5ml of sodium carbonate (Na_2CO_3) solution. To this add 22.5ml of sodium bicarbonate solution and made upto 100ml with distilled water which corresponds to 0.2 M sodium carbonate and bicarbonate buffer.

Standardise pH meter and measure the pH of required buffer. This gives the Carbonate- bicarbonate buffer pH 10.2.

RESULT:

Carbonate bicarbonate buffer was prepared and pH observed was 7.5 which was adjusted to 10.2 using 1N Hcl and 5N NaoH.

5. PHOSPHATE BUFFER:**REAGENTS REQUIRED:**

- Monobasic: Dissolve 2.78gm of sodium dihydrogen phosphate in 100ml of distilled water.
- Dibasic sodium phosphate (0.2M): Dissolve 5.3gm of disodium hydrogen phosphate or 7.17 gm sodium hydrogen phosphate in 100ml distilled water.

PROCEDURE:

39 ml of dihydrogen sodium phosphate is mixed with 61 ml of disodium hydrogen phosphate This made up to 200ml with distilled water .This gives phosphate (Po_4)₂ buffer of 0.2M.

Standardized pH meter with standard buffer. Washed electrode with distilled water and introduced it into phosphate buffer prepared. The pH of the solution is 6.8.

RESULT:

Phosphate buffer was prepared and pH was observed 8.5 which was made upto 6.8 using 1N Hcl and 5N NaoH.

6. POTASSIUM PHOSPHATE BUFFER:**REAGENTS REQUIRED:**

- Dipotassium hydrogen phosphate
- Potassium dihydrogen phosphate

PROCEDURE:

174.18 g/mol dipotassium hydrogen phosphate and 136.09 g/mol potassium dihydrogen phosphate was taken and made up to 200ml using distilled water. This gives the potassium buffer.

Standardised pH meter with standard buffer. Washed electrode with distilled water and introduced it into potassium buffer prepared. The pH of the solution is 6.5.

RESULT:

Dipotassium hydrogen phosphate (K_2HPO_4) and potassium dihydrogen phosphate (KH_2PO_4) solution were prepared and the pH was measured to be 9.87 and 4.23 respectively, the solution were made using 1N HCl and 5N NaOH respectively and the pH was found to be 6.5.