

Calcium carbonate

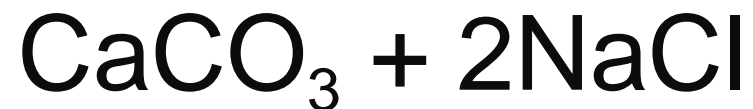
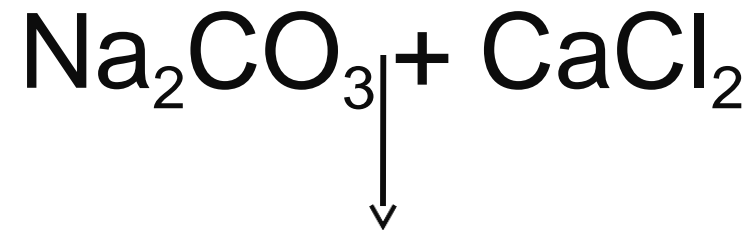
CaCO_3 / 100.89

Synonym:

**Precipitated chalk,
Precipitated calcium
carbonate**

Method Of Preparation:

Prepared from sodium carbonate
and calcium chloride.

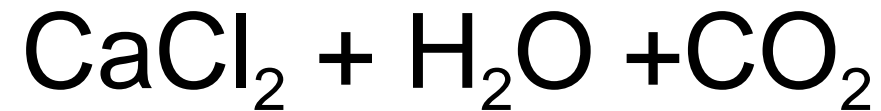
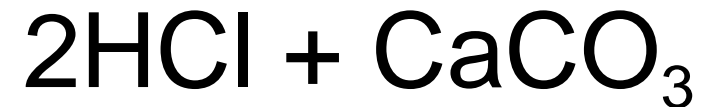


Physical Properties:

- White powder
- Tasteless
- Odorless
- Practically insoluble in water
- Insoluble in Ethanol
- Stable in air

b) Chemical Properties:-

When it react with hydrochloric acid gives calcium chloride, water



Uses:

- Gastrointestinal agent
- Antacid
- Electrolyte replinisher

Storage

“ It should be stored in well closed container at a cool Place.”

Magnesium Carbonate



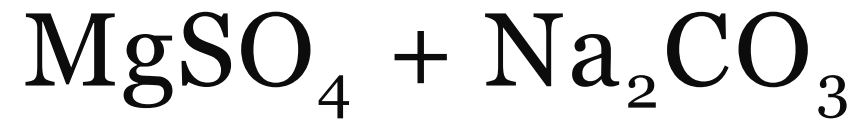
2 Types:-differ in bulk density (15 gm)

☐ Heavy magnesium carbonate (30 ml)

☐ Light magnesium carbonate (125ml)

Method of preparation :

By reacting magnesium sulphate
& sodium carbonate

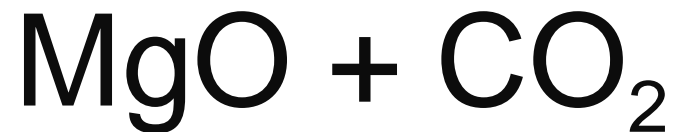
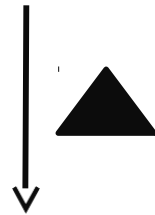


Physical Properties:

- White powder
- Tasteless
- Odorless
- Practically insoluble in water
- Insoluble in Ethanol
- Soluble in dilute mineral acids

b) Chemical Properties:-

On heating convert in magnesium oxide.



Uses:

- Gastrointestinal agent
- Antacid
- Used in tooth powder
- Used in cosmetics
- Used in silver Polishes

Storage

“ It should be stored in well closed container at a cool Place.”