

AIM: To prepare and standardize 0.1 N NaOH.

APPARATUS: burette, pipette, conical flask etc.

CHEMICALS: 0.1 N NaOH, 0.1 N oxalic acid, phenolphthalein.

PRINCIPLE: Standardisation of sodium hydroxide is an example of acidimetric titration. When oxalic acid is allowed to react with sodium hydroxide, sodium oxalate and water are obtained. This is an example of acid base reaction. In this titration, to locate the end point phenolphthalein indicator is used

PROCEDURE: Into a conical flask pipette out exactly 10 ml of 0.1 N oxalic acid and add one or two drops of phenolphthalein indicator. Titrate the solution against sodium hydroxide solution until a permanent pink color is obtained. Repeat the titration to get concurrent values.

OBSERVATION TABLE:

S.No.	Volume of 0.1 N oxalic acid (ml)	Burette reading		Volume of NaOH used (ml)
		Initial (ml)	Final (ml)	

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FORMULA: The normality of NaOH is calculated using the formula

$$N_1V_1=N_2V_2$$

RESULT: The normality of the given sodium hydroxide solution is